

immersing the plating barrel and the graphite particles contained therein in an aqueous electrolyte media comprising an acid and an oxidizing agent; subjecting the immersed graphite particles to an anodic current; rotating the plating barrel; removing the graphite particles from the electrolyte and rinsing the graphite particles with a solvent; and removing the excess electrolyte and solvent from the graphite particles.

2(original). The method of Claim 1 wherein the electrolyte comprises  $\text{H}_2\text{SO}_4$  for the acid and  $\text{HNO}_3$  for the oxidizing agent.

3(original). The method of Claim 2 wherein the electrolyte comprises between approximately 99 Vol.% and 50 Vol.% of 66 Wt.%  $\text{H}_2\text{SO}_4$  and between approximately 1 Vol.% and 50 Vol.% of 40 Wt.%  $\text{HNO}_3$ .

4(original). The method of Claim 1 or Claim 3 wherein the density of the current to which the immersed graphite particles are subjected is between approximately 0.1 mA and 5A per gram of graphite.

5(original). The method of Claim 4 wherein the immersed graphite particles are subjected to the anodic current for between approximately 1 minute and 180 minutes.

6(original). The method of Claim 3 or 5 wherein the solvent is distilled water.

7(original). The method of Claim 3 or 5 wherein the solvent is deionized water.

8(original). The method of Claim 1 wherein the acid is  $\text{H}_2\text{SO}_4$  and the oxidizing agent is selected from the group consisting of  $\text{HNO}_3$ ,  $\text{CrO}_3$ ,  $\text{KMnO}_4$ ,  $(\text{NH}_4)_2\text{SO}_4$ ,  $\text{PbO}_2$ ,  $\text{MnO}_2$ ,  $\text{MnO}$ ,  $\text{H}_2\text{O}_2$  and  $\text{HClO}_4$ .

9(previously presented). A method of preparing a graphite intercalation compound comprising:

providing graphite particles;

providing a plating barrel;

placing the graphite particles in the plating barrel;

immersing the plating barrel and the graphite particles

contained therein in an aqueous electrolyte media of

between about 90 vol.% and 75 Vol.% of 66 Wt.%  $\text{H}_2\text{SO}_4$  and

between about 10 Vol.% and 25 Vol.% of 40 Wt.%  $\text{HNO}_3$ ;

subjecting the immersed graphite particles to a current of

approximately 1mA per gram of graphite for between

approximately 1 to 60 minutes;

rotating the plating barrel;

rinsing the graphite particles in water for approximately 1 minute; and

drying the graphite particles.

10(original). The method of Claim 9 wherein the water is distilled.

11(original). The method of Claim 9 wherein the water is deionized.

12(original). The method of Claim 9 wherein the graphite particles are dried in a vacuum drier.

13(original). The method of Claim 9 wherein the graphite particles are dried in a filter press.

14(original). The method of Claim 9 wherein the graphite particles are dried in a centrifuge.

15(canceled).

16(previously presented). The method of Claim 1 or 9 wherein the plating barrel is a wall plating barrel with a +50 mesh opening wall.

17(canceled).

18(canceled).

19(original). The method of Claims 1 or 9 wherein the graphite particles are selected from the group consisting of natural, synthetic, vein, and amorphous graphite, all having a purity of between about 80% and 99.9% LOI.